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continued 9. Describe the
typical physical properties
of substances that have

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Bonds 1. A covalent bond forms when two or more valence electrons are attracted by the positively charged nuclei of two atoms and thus are shared between both atoms. 2. The H₂ molecule is stable because

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each hydrogen atom now has a shared pair of electrons and has achieved a stable noble gas configuration. 3.

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Bonds Compounds Concept
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is a unit used to measure
the number of atoms,
molecules, or (in the case
of ionic compounds) formula
units in a given mass of a
substance.

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Complete each statement below by choosing a term from the following list. ...

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answers chapter 4. stances.

2. Answers will vary but may include iron rusting, milk.

. .

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The mole is a unit used to

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Compounds Concept
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measure the number of atoms, molecules, or (in the case of ionic compounds) formula units in a given mass of a substance. The mole is defined as the amount of substance that contains the number of carbon atoms in

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exactly 12 g of carbon-12,
Avogadro's number (6.022×10^{23}) of atoms of
carbon-12.

3.2: The Mole Concept and
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Holt Chemistry 7 Covalent
Compounds Name Class Date
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Concept Review continued Use
the system of prefixes and
the rule for suffixes to
name the following
compounds. 15. PbCl_2 16.

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Complete each statement
below by choosing a term
from the ...

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force between any two
particles is the ionic bond,

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in which two ions . . .

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Ions and Ionic Compounds

... Holt Chemistry 19 Ions and

Ionic Compounds For a

substance to conduct an

electric current, the

substance must possess free-

moving charged particles.

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These charged particles may be delocalized electrons, such as those found in substances that form metallic bonds.

Holt Chemistry Ionic Bonding
Salts Answers

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1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the electron-proton attraction

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is stronger than the
electron-electron and proton-
proton repulsions. The atoms
are drawn to each other and
their potential energy
decreases.

6 Chemical Bonding

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A covalent bond, where the electrons are shared equally is called a nonpolar bond (eg H-H) and an unequal sharing of the pair of bonding electrons results in

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a polar bond. The unequal sharing of electrons is due to the ability of an atom to attract electrons towards itself which is known as Electronegativity.

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bonding theory that is based
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radical, or molecule that
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an ion that has a positive

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charge due to losing
electrons an ion that has a
negative charge due to
gaining electrons

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a concept of chemical bonding theory that is based on the assumption that atoms tend to have either empty valence shells or full valence shells of eight electrons. Ion, an atom, radical, or molecule that

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has gained or lost one or more electrons and has a negative or positive charge. cation.

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