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Reactions (CdCl_2) and potassium sulfide

(K_2S), are poured together. As the solutions

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mix, a bright yellow solid forms. Name The type and condition of the ground surface during non-growing periods dictates the range for ET_c . Where the surface is bare soil, then the K_c will be quite similar to the K_c ini predicted in Chapter

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Reactions Keywords: chapter,11,experiment

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Experiment 17 Precipitation Reactions 109
Name Date Class Set 3. 0.1M barium
chloride, BaCl_2 . 0.1M magnesium nitrate,

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$\text{Mg}(\text{NO}_3)_2$. 0.2M sodium chromate,
 Na_2CrO_4 . 0.2M aluminum sulfate,
 $\text{Al}_2(\text{SO}_4)_3$. 0.2M potassium chromate,
 K_2CrO_4 . 0.2M silver nitrate, AgNO_3 .
SAFETY FIRST! In this lab, the solutions
you use may contain harmful materials.

Name

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The aim of this part of the experiment is to determine the factors that control whether or not a precipitate forms, i.e. is it the sum of the ion concentrations (in this case $[Ag^+]$

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+ [Cl⁻]) or the product of the ion concentrations (in this

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Aqueous silver nitrate (AgNO₃) is added to a solution containing potassium chloride (KCl), and the precipitation of a white solid,

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silver chloride (AgCl), is observed: AgNO_3
(aq) + KCl (aq) \rightarrow AgCl (s) + KNO₃ (aq)

Note that the product silver chloride is the precipitate, and it is designated as a solid.

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This chapter provides an overview of downstream processing. After completing this chapter, students will be able to: describe the general elements of the downstream purification process for a

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typical monoclonal antibody-based
therapeutic, starting with the bioreactor
harvest pool and finishing with pre-
formulation, pure drug product.

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