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4/10 Short

Answers 1) (a)

In the reaction

$2\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_3\text{O}^+$

$+\text{OH}^-(aq)$

use the Bronsted-  
Lowry Theory to  
discuss the

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Acidic and/or  
basic nature of  
water. [2]  $\text{H}_2\text{O}$   
 $+\text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+$   
 $+\text{OH}^-$  Water can  
donate a proton  
(BL acid) or  
receive a proton  
(BL base)

**ANSWERS Unit 14:  
Review Acids and  
Bases**

acids react with

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bases in a  
neutralization  
reaction( also  
double  
replacement).  
applications:  
the reaction of  
an acid and a  
base produces a  
salt and water.  
List the common  
strong acids and  
the strong  
bases.

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Name 3

characteristics  
that acids and  
bases share.

Acids and bases  
can both be .

corrosive.

poisonous.

electrolytes

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(can conduct  
electricity)  
Guide Answers

Both can also be described as having unique tastes and capable of changing the color of indicators.

Describe the process of dissociation of phosphoric acid

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( $\text{H}_3\text{PO}_4$ ). During the dissociation of acid, the solute ( $\text{H}_3\text{PO}_4$ ) is dissolved in water.

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$$10^{-4} = 5.25 \times$$

$$10^{-4x} = 8.$$

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Answers. 1)

Using your

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knowledge of the  
Brønsted-Lowry  
theory of acids  
and bases, write  
equations for  
the following  
acid-base  
reactions and  
indicate each  
conjugate acid-  
base pair: a)  
 $\text{HNO}_3 + \text{OH}^- \rightarrow$   
 $\text{H}_2\text{O} + \text{NO}_3^-$ .  
 $\text{HNO}_3$  and  $\text{NO}_3^-$

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And Bases pair

$\text{OH}^-$  and  $\text{H}_2\text{O}$

make the other.

b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O}$

(  $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$

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our-

Explanation-1.

2. Bases react

with. acids to

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And Bases  
produce salts  
and water. water  
to

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A Base is any  
substance whose  
aqueous solution  
is characterized  
by a bitter  
taste, a

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slippery feel,  
the ability to  
turn red litmus  
blue, and the  
ability to react  
with acids to  
form salts.  
Questions...

## **Answers about Acids and Bases**

Solution: The  
reaction of the  
acid + base

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gives a product of salt + water, which is considered as neutralization reaction.

Examples:  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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Chapter 2 Acid  
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indicate the  
conjugate acid-  
base pairs. base

2 CO<sub>3</sub><sup>2-</sup> acid 2

HCO<sub>3</sub><sup>-</sup> base 1 NH<sub>3</sub>

3 True b. A

proton has been  
transferred from  
acid 1 to base 2

in the above

reaction. True

or False? 3.

Consider the

neutralization

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And Bases

described by the  
equation:  $\text{HCO}_3(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l})$

3(aq) OH (aq) ??

CO<sub>3</sub><sup>2-</sup> (aq) H<sub>2</sub>O

20(1) a. Label

the conjugate

acid-base pairs

in this system.

HCO<sub>3</sub><sup>-</sup>

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ammonia and  
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upcoming test.

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acid -

substances that  
ionize in  
solutions to  
form  $H^+$  ions.

amphoteric - a  
substance that  
can be an acid

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Arrhenius Model  
- in aqueous solutions, acids form hydrogen ions ( $H^+$ ). base - substances that ionize in solutions and form  $OH^-$  ions. binary acids - acids that do not contain oxygen in their

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